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· These structures are helpful in explanation of chemical reactivity or the chemical reaction of the compound thats why we can say resonance phenomenon is the result of mesomeric effect or delocalisation. (i) Resonating structure are not the real structures of conjugated compounds. (iii) The real structure of conjugated compound is a hybrid of all resonating structures. This phenomenon is known as resonance, mesomerism or delocalisation.

(iii) Thus resonance is nothing but hybridisation of resonating structures and resonance phenomenon will (iv) Conditions of Resonating Structures : Resonance structures should fulfil following conditions : (a) All resonating structures must have the same arrangement of atomic nuclei. Resonance differs from tautomerism in this very important aspect.

Position of hydrogen nuclei in (I) and (II) are different, hence (I) and (II) are not resonating structures, they

(b) The resonating structures must have the same numbers of paired and unpaired electrons. However, they differ in the way of distribution of electrons.

 All atoms that are part of the delocalisation system must be in a plane or be nearly planar. (v) All atoms of the resonating structure should follow the octet rule. For Example: All atoms follow octet rule.

$$\bigcap_{I}^{\otimes H_{I}} \bigoplus_{I}^{NH_{I}} \longrightarrow \bigcap_{I}^{NH_{I}} \longrightarrow \bigcap_{I}$$

1. If there are two π bonds at alternate position then e^{ϵ} of one π bond are transferred towards another π bond. (According to I-effect).

JEE-Mathematics

7. VALUES OF T-RATIOS OF SOME STANDARD ANGLES:

Angles T-ratio	0	30 π/6	45 π/4	60 π/3	90 π/2	180 π	270 3π/2
	0						
$\sin\theta$	0	1/2	$1/\sqrt{2}$	√3/2	1	0	-1
$\cos\theta$	1	$\sqrt{3}/2$	$1/\sqrt{2}$	1/2	0	-1	0
tan 0	0	1/√3	1	√3	N.D.	0	N.D.
cot θ	N.D.	√3	1	$1/\sqrt{3}$	0	N.D.	0
sec θ	1	2/√3	$\sqrt{2}$	2	N.D.	-1	N.D.
cosecθ	N.D.	2	$\sqrt{2}$	2/√3	1	N.D.	-1

N.D. → Not Defined

 $\sin n\pi = 0$; $\cos n\pi = (-1)^n$; $\tan n\pi = 0$ where $n \in I$

(b) $\sin(2n+1)\frac{\pi}{2} = (-1)^n; \cos(2n+1)\frac{\pi}{2} = 0 \text{ where } n \in I$

Illustration 4 : If
$$\sin \theta = -\frac{1}{2}$$
 and $\tan \theta = \frac{1}{\sqrt{3}}$ then θ is equal to -

(A) 30 (B) 150 (C) 210 Let us first find out
$$\theta$$
 lying between 0 and 360 .

(D) none of these

Since
$$\sin \theta = -\frac{1}{2} \Rightarrow \theta = 210$$
 or 330 and $\tan \theta = \frac{1}{\sqrt{3}} \Rightarrow \theta = 30$ or 210

Hence , $\theta = 210$ or $\frac{7\pi}{6}$ is the value satisfying both.

Ans. (C)

Do yourself - 3:

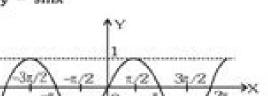
Solution :

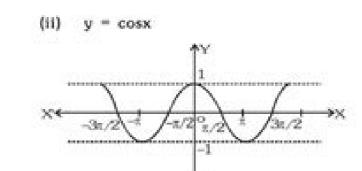
(i) If
$$\cos\theta = -\frac{1}{2}$$
 and $\pi < \theta < \frac{3\pi}{2}$, then find the value of $4\tan^2\theta - 3\csc^2\theta$.

(b)
$$\tan \frac{11\pi}{3} - 2\sin \frac{9\pi}{3} - \frac{3}{4}\csc^2 \frac{\pi}{4} + 4\cos^2 \frac{17\pi}{6} = \frac{3 - 2\sqrt{3}}{2}$$

GRAPH OF TRIGONOMETRIC FUNCTIONS:

(i) y = sinx





(IV) Rearranement reaction : It may be defined as the reaction involving reshuffling of the sequence of atom to form a new structure :

In this reaction ketoxime reacts with $H_2SO_4(conc.)$ to give a product by alkyl migration.

 Bayer-villiger rearrangement : The reaction involve the oxidation of ketone to ester by the treatment with peracids such as peracetic acid, perbenzoic acid, pertrifluro acetic acid, etc. In this case best migratory group migrate to electron deficient

· Acid-catalysted decomposition of cumene hydroxide

$$\bigcirc \overset{\circ}{\circ} \overset{$$

THE s-BLOCK ELEMENTS 301

Table 10.1 Atomic and Physical Properties of the Alkali Metals

Property	Lithium Li	Sodium Na	Potassium K	Rubidium Rb	Caesium Cs	Francium Fr
Atomic number	3	11	19	37	55	87
Atomic mass (g mol ⁻¹)	6.94	22.99	39.10	85.47	132.91	(223)
Electronic configuration	[He] 2s1	[Ne] 3s ¹	[Ar] 4s ¹	[Kr] 5s ¹	[Xe] 6s ¹	[Rn] 7s ¹
Ionization enthalpy / kJ mol ⁻¹	520	496	419	403	376	~375
Hydration enthalpy/kJ mol ⁻¹	-506	-406	-330	-310	-276	1
Metallic radius / pm	152	186	227	248	265	,0
Ionic radius M* / pm	76	102	138	152	167	(180)
m.p. / K	454	371	336	312	302	((=)
b.p / K	1615	1156	1032	961	944	(Ce)
Density / g cm ⁻³	0.53	0.97	0.86	1.53	1.90	((=)
Standard potentials E°/V for (M*/M)	-3.04	-2.714	-2.925	-2.930	-2.927	15.00
Occurrence in lithosphere [†]	18*	2.27**	1.84**	78-12*	2-6*	~ 10 ⁻¹⁸ *

*ppm (part per million), ** percentage by weight; † Lithosphere: The Earth's outer layer: its crust and part of the upper mantle

This property makes caesium and potassium useful as electrodes in photoelectric cells.

10.1.6 Chemical Properties

The alkali metals are highly reactive due to their large size and low ionization enthalpy. The reactivity of these metals increases down the group.

(i) Reactivity towards air: The alkali metals tarnish in dry air due to the formation of their oxides which in turn react with moisture to form hydroxides. They burn vigorously in oxygen forming oxides. Lithium forms monoxide, sodium forms peroxide, the other metals form superoxides. The superoxide O₂ ion is stable only in the presence of large cations such as K, Rb, Cs.

 $4 \text{ Li } + \text{O}_2 \rightarrow 2 \text{ Li}_2\text{O} \text{ (oxide)}$

 $2 \text{ Na} + \text{O}_2 \rightarrow \text{Na}_2 \text{ O}_3 \text{ (peroxide)}$

 $M + O_2 \rightarrow MO_2$ (superoxide)

(M = K, Rb, Cs)

In all these oxides the oxidation state of the alkali metal is +1. Lithium shows exceptional behaviour in reacting directly with nitrogen of air to form the nitride, Li₃N as well. Because of their high reactivity towards air and water, alkali metals are normally kept in kerosene oil.

Problem 10.1

What is the oxidation state of K in KO₂?

Solution

The superoxide species is represented as O_2^- ; since the compound is neutral, therefore, the oxidation state of potassium is ± 1 .

2019-20

Subnetting Cheat Sheet

CIDR	SUBNET MASK	ADDRESSES	WILDCARD MASK	Classful IPv4 Addresses		
/32	255.255.255. 255	1	0.0.0.0	Class A	0.0.0.0 - 127.255.255.255	
/31	255.255.255. 254	2	0.0.0.1	Class 8	128.0.0.0 - 191.255.255.255	
/30	255.255.255. 252	4	0.0.0.3	Class C	192.0.0.0 - 223.255.255.255	
/29	255.255.255.248	8	0.0.0.7	Class D	224.0.0.0 - 239.255.255.255	
/28	255.255.255. 240	16	0.0.0.15	Class E	240.0.0.0 - 255.255.255.255	
/27	255.255.255. 224	32	0.0.0.31			
/26	255.255.255.192	64	0.0.0.63	Private IPv4 Addresses		
/25	255.255.255.128	128	0.0.0.127	10.0.0.0 - 10.255.255.255		
/24	255.255.255.0	256	0.0.0.255	172.16.	i.0.0 - 172.31.255.255	
/23	255.255.254.0	512	0.0.1.255	192.168.0.0 - 192.168.255.255		
/22	255.255.252.0	1024	0.0.3.255	Special IPv4 Addresses		
/21	255.255.248.0	2048	0.0.7.255			
/20	255.255.240.0	4096	0.0.15.255	Local Host	127.0.0.0 - 127.255.255.255	
/19	255.255.224.0	8192	0.0.31,255	APIPA	169.254.0.0 - 169.254.255.255	
/10	255.255.192.0	16384	0.0.63.255	Bogon IPv4 Addresses		
/17	255.255.128.0	32768	0.0.127.255			
/16	255.255.0.0	65536	0.0.255.255	0.0.0.0/8	This network	
/15	255.254.0.0	131072	0.1.255.255	10.0.0.0/8	Private IPv4 Address Block	
/14	255.252.0.0	262144	0.3.255.255	100.64.0.0/10	Carrier-grade NAT	
/13	255.248.0.0	524288	0.7.255.255	127.0.0.0/8	Loopback	
/12	255.240.0.0	1048576	0.15.255.255	127.0.53.53	Name collision occurrence	
/11	255.224.0.0	2097152	0.31.255.255	169.254.0.0/16	Link local	
/10	255.192.0.0	4194304	0.63.255.255	172.16.0.0/12	Private IPv4 Address Block	
/9	255.128.0.0	8388608	0.127.255.255	192.0.0.0/24	IETF protocol assignments	
/8	255.0.0.0	16777216	0.255.255.255	192.0.2.0/24	TEST-NET-1	
/7	254.0.0.0	33554432	1.255.255.255	192.168.0.0/16	Private IPv4 Address Block	
/6	252.0.0.0	67108864	3 .255.255.255	198.18.0.0/15	Network benchmark testing	
/5	248.0.0.0	134217728	7 .255.255.255	198.51.100.0/24	TEST-NET-2	
/4	240.0.0.0	268435456	15. 255.255.255	203.0.113.0/24	TEST-NET-3	
/3	224.0.0.0	536870912	31. 255.255.255	224.0.0.0/4	Multicast	
/2	192.0.0.0	1073741824	63. 255.255.255	240.0.0.0/4	Reserved	
/1	128.0.0.0	2147483648	127. 255.255.255	255.255.255.255/32		
/0	0.0.0.0	4294967296	255. 255.255.255	www.ipcisco.com		



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